

Name: _____

Period: _____

Seat#: _____

Directions: Show all work in a way that would earn you credit on the AP Test! This is always the rule! Some answers are provided at the end in italics and underlined. If you need more space, use binder paper and staple to your worksheet.

NChO 1999

Reaction	ΔH
$\text{Mg(s)} + 2 \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$	-467 kJ mol^{-1}
$\text{MgO(s)} + 2 \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{O(l)}$	-151 kJ mol^{-1}

27. According to this information, and given the fact that for water, $\Delta H_f = -286 \text{ kJ mol}^{-1}$, what is ΔH_f for MgO(s) ?
- (A) -904 kJ mol^{-1} (C) -334 kJ mol^{-1}
 (B) -602 kJ mol^{-1} (D) -30 kJ mol^{-1}

NChO 1998

22. Carbon reacts with oxygen according to this equation. $2\text{C(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{CO(g)}$ $\Delta H = -220 \text{ kJ}$

Which statements are true?

- The reaction is exothermic.
- The combustion of 0.50 mol of carbon produces 55 kJ of heat energy

- (A) 1 only (C) both 1 and 2
 (B) 2 only (D) neither 1 nor 2

24. Use these data to calculate ΔH° for this reaction. $\text{NO(g)} + \frac{1}{2} \text{O}_2\text{(g)} \rightarrow \text{NO}_2\text{(g)}$

- (A) -57.0 (C) $+28.5$
 (B) -28.5 (D) $+57.0$

Reaction	$\Delta H^\circ, \text{ kJ mol}^{-1}$
$\frac{1}{2} \text{N}_2\text{(g)} + \frac{1}{2} \text{O}_2\text{(g)} \rightarrow \text{NO(g)}$	90.2 kJ mol^{-1}
$\frac{1}{2} \text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow \text{NO}_2\text{(g)}$	33.2 kJ mol^{-1}

25. A 1.0 g sample of substance A at 100°C is added to 100 ml of H_2O at 25°C . Using separate 100 mL portions of H_2O , the procedure is repeated with substance B and then with substance C. How will the final temperatures of the water compare?

- (A) $T_c > T_b > T_a$ (C) $T_a > T_b > T_c$
 (B) $T_b > T_a > T_c$ (D) $T_a = T_b = T_c$

Substance	Specific Heat
A	$0.60 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$
B	$0.40 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$
C	$0.20 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

26. How many grams of benzene, $\text{C}_6\text{H}_6\text{(l)}$, must be burned in a bomb calorimeter to raise its temperature by 1.5°C ? Given: The calorimeter constant is $12.59 \text{ kJ }^\circ\text{C}^{-1}$ and the $\Delta H^\circ_{\text{combustion}}$ for $\text{C}_6\text{H}_6 = -41.9 \text{ kJ g}^{-1}$
- (A) 0.45 g (C) 3.3 g
 (B) 2.8 g (D) 8.4 g

NChO 1997

19. 30.0 mL of water at $10.^\circ\text{C}$ is mixed with 50.0 mL of water at $60.^\circ\text{C}$. What is the final temperature of the mixture?

- (A) 31°C (C) 41°C
 (B) 35°C (D) 46°C

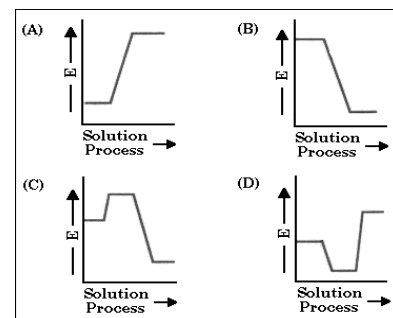
24. What is the value of ΔH° for this reaction? $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{O(l)} \rightarrow 2\text{Fe(OH)}_3\text{(s)}$

- (A) 35.6 kJ (C) 858.6 kJ
 (B) 286 kJ (D) -536 kJ

Substance	$\Delta H^\circ_f \text{ (kJ mol}^{-1}\text{)}$
$\text{Fe}_2\text{O}_3\text{(s)}$	-824.2
$\text{Fe(OH)}_3\text{(s)}$	-823.0
$\text{H}_2\text{O(l)}$	-285.8

**Dougherty Valley HS Chemistry - AP
Thermochemistry – NChO**

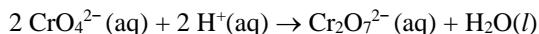
25. When $\text{Na}_2\text{S}_2\text{O}_3 \cdot 3\text{H}_2\text{O}$ dissolves in water, the solution gets cold. Which energy diagram best represents the behavior of this solution process?



NChO 1996

22. The standard enthalpy of formation (ΔH°_f) for sodium bromide is the enthalpy change for the reaction
- (A) $\text{Na}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{NaBr}(\text{g})$
 (B) $\text{Na}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{NaBr}(\text{s})$
 (C) $2 \text{Na}(\text{s}) + \text{Br}_2(\text{g}) \rightarrow 2 \text{NaBr}(\text{s})$
 (D) $\text{Na}(\text{s}) + \frac{1}{2} \text{Br}_2(\text{l}) \rightarrow \text{NaBr}(\text{s})$

23. Use the standard enthalpies of formation to calculate ΔH° for this reaction:



- (A) 272.1 kJ (C) -13.7 kJ
 (B) 13.7 kJ (D) -272.1 kJ

Substance	$\Delta H^\circ_f, \text{kJ mol}^{-1}$
$\text{CrO}_4^{2-}(\text{aq})$	- 881.2
$\text{Cr}_2\text{O}_7^{2-}(\text{aq})$	- 1490.3
$\text{H}^+(\text{aq})$	0
$\text{H}_2\text{O}(\text{l})$	- 285.8

NChO 1995

21. For which of these processes is the sign of the enthalpy change different from the others?

- (A) $\text{Al}_2\text{O}_3(\text{s}) \rightarrow 2 \text{Al}(\text{s}) + 3/2 \text{O}_2(\text{g})$
 (B) $\text{H}_2\text{O}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l})$
 (C) $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$
 (D) $\text{Cl}(\text{g}) + \text{e}^- \rightarrow \text{Cl}^-(\text{g})$

22. The standard enthalpy of formation (ΔH°) for nitrogen(IV) oxide is the enthalpy change for the reaction

- (A) $\text{N}(\text{g}) + 2\text{O}(\text{g}) \rightarrow \text{NO}_2(\text{g})$
 (B) $\frac{1}{2} \text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2(\text{g})$
 (C) $\frac{1}{2} \text{N}_2\text{O}_4(\text{g}) \rightarrow \text{NO}_2(\text{g})$
 (D) $\text{NO}(\text{g}) + \frac{1}{2} \text{O}_2 \rightarrow \text{NO}_2(\text{g})$

23. In a bomb calorimeter, reactions are carried out at

- (A) constant volume. (C) 1 atm pressure and 25 °C.
 (B) constant pressure. (D) 1 atm pressure and 0 °C.

26. Consider the reaction $\text{Hg}(\text{l}) + 2 \text{Ag}^+(\text{aq}) \rightarrow \text{Hg}^{2+}(\text{aq}) + 2 \text{Ag}(\text{s})$

What is enthalpy change for this rxn if ΔH°_f for $\text{Ag}^+(\text{aq})$ is $+105.6 \text{ kJ mol}^{-1}$ and for $\text{Hg}^{2+}(\text{aq})$ is $+171.1 \text{ kJ mol}^{-1}$?

- (A) 65.5 kJ are evolved per mole of Hg.
 (B) 65.5 kJ are absorbed per mole of Hg.
 (C) 40.1 kJ are evolved per mole of Hg.
 (D) 40.1 kJ are absorbed per mole of Hg.

NChO 1994

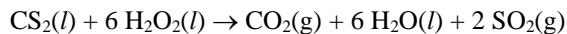
24. A student mixes 100 mL of 0.50 M NaOH with 100 mL of 0.50 M HCl in a styrofoam cup and observes a temperature increase of ΔT_1 . When she repeats this experiment using 200mL of each solution, she observes a temperature change of ΔT_2 . If no heat is lost to the surroundings or absorbed by the styrofoam cup, what is true about ΔT_1 and ΔT_2 ?

- (A) $\Delta T_2 = \Delta T_1$ (C) $\Delta T_2 = 2 \Delta T_1$
 (B) $\Delta T_2 = 0.5 \Delta T_1$ (D) $\Delta T_2 = 4 \Delta T_1$

**Dougherty Valley HS Chemistry - AP
Thermochemistry – NChO**

27. Given these values of ΔH° :
- | | |
|--|-------------------------------------|
| $\text{CS}_2(l) + 3\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{SO}_2(g)$ | $\Delta H^\circ = -1077 \text{ kJ}$ |
| $\text{H}_2(g) + \text{O}_2(g) \rightarrow \text{H}_2\text{O}_2(l)$ | $\Delta H^\circ = -188 \text{ kJ}$ |
| $\text{H}_2(g) + (1/2) \text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$ | $\Delta H^\circ = -286 \text{ kJ}$ |

What is the value of ΔH° for this reaction?



- (A) -1175 kJ (C) -1665 kJ
(B) -1551 kJ (D) -3921 kJ

NChO 1993

13. Which process or reaction has a positive ΔH ?
- (A) $\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(s)$
(B) $2\text{CH}_3\text{OH}(l) + 3\text{O}_2(g) \rightarrow 2\text{CO}_2(g) + 4\text{H}_2\text{O}(l)$
(C) $\text{CO}_2(s) \rightarrow \text{CO}_2(g)$
(D) $2 \text{Na}(s) + \text{Cl}_2(g) \rightarrow 2 \text{NaCl}(s)$
15. For the reaction $\text{H}_2(g) + \text{I}_2(s) \rightarrow 2 \text{HI}(g)$ $\Delta H_{\text{rxn}} = 53.0 \text{ kJ}$
What will be the value of ΔH_{rxn} (in kJ) for this rxn ?
 $\text{HI}(g) \rightarrow \frac{1}{2} \text{H}_2(g) + \frac{1}{2} \text{I}_2(s)$
- (A) 26.5 (C) -26.5
(B) 7.3 (D) -53.0

Answers

- | | |
|-------------|-------------|
| <u>1999</u> | <u>1995</u> |
| 27) B | 21) D |
| | 22) B |
| <u>1998</u> | 23) A |
| 22) C | 26) C |
| 24) A | |
| 25) C | <u>1994</u> |
| 26) A | 24) A |
| | 25) 27 |
| <u>1997</u> | |
| 19) C | <u>1993</u> |
| 24) A | 13) C |
| 25) A | 15) C |
| <u>1996</u> | |
| 22) D | |
| 23) C | |