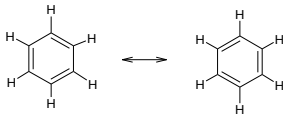
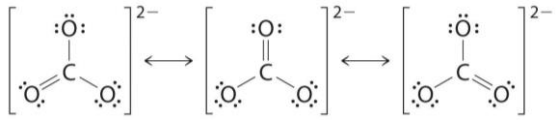
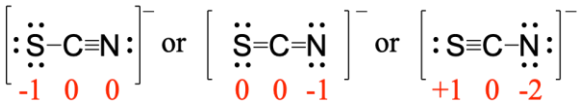


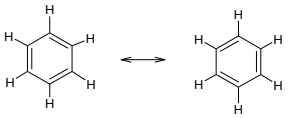


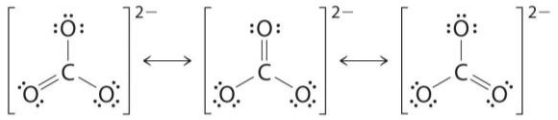


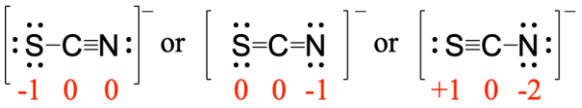
X + E	Overall Structure (Electronic Geometry)	Forms
2	Linear	AX_2
3	Trigonal Planar	AX_3, AX_2E
4	Tetrahedral	AX_4, AX_3E, AX_2E_2
5	Trigonal bipyramidal	AX_5, AX_4E, AX_3E_2, AX_2E_3
6	Octahedral	AX_6, AX_5E, AX_4E_2

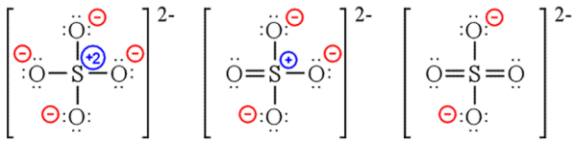
"Rules" for Drawing Lewis Structures

1) Count and sum valence electrons.
2) Place your atoms.
3) Bond all atoms w/ a single bond (try simplest way 1st).
4) Give all atoms a full shell.
5) Re-count the electrons you used.
6) Used too few? Put extras on the central atom.
7) Used too many? Try double or triple bonds to fix.
8) Re-count.





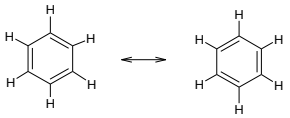


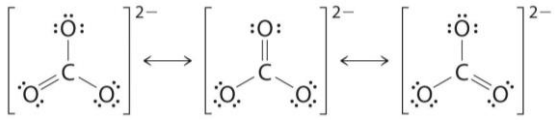


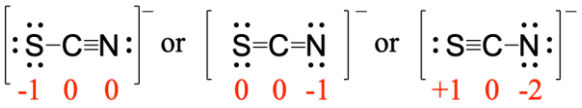
X + E	Overall Structure (Electronic Geometry)	Forms
2	Linear	AX_2
3	Trigonal Planar	AX_3, AX_2E
4	Tetrahedral	AX_4, AX_3E, AX_2E_2
5	Trigonal bipyramidal	AX_5, AX_4E, AX_3E_2, AX_2E_3
6	Octahedral	AX_6, AX_5E, AX_4E_2

"Rules" for Drawing Lewis Structures

1) Count and sum valence electrons.
2) Place your atoms.
3) Bond all atoms w/ a single bond (try simplest way 1st).
4) Give all atoms a full shell.
5) Re-count the electrons you used.
6) Used too few? Put extras on the central atom.
7) Used too many? Try double or triple bonds to fix.
8) Re-count.







X + E	Overall Structure (Electronic Geometry)	Forms
2	Linear	AX_2
3	Trigonal Planar	AX_3, AX_2E
4	Tetrahedral	AX_4, AX_3E, AX_2E_2
5	Trigonal bipyramidal	AX_5, AX_4E, AX_3E_2, AX_2E_3
6	Octahedral	AX_6, AX_5E, AX_4E_2

"Rules" for Drawing Lewis Structures

1) Count and sum valence electrons.
2) Place your atoms.
3) Bond all atoms w/ a single bond (try simplest way 1st).
4) Give all atoms a full shell.
5) Re-count the electrons you used.
6) Used too few? Put extras on the central atom.
7) Used too many? Try double or triple bonds to fix.
8) Re-count.