

Name: _____

Period: _____

Seat#: _____

Directions: Any worksheet that is labeled with an * means it is suggested extra practice. We do not always have time to assign every possible worksheet that would be good practice for you to do. You can do this worksheet when you have extra time, when you finish something early, or to help you study for a quiz or a test. If and when you choose to do this Extra Practice worksheet, please do the work on binder paper. You will include this paper stapled into your Rainbow Packet when you turn it in, even if you didn't do any of this. We want to make sure we keep it where it belongs so you can do it later if you want to (or need to). If you did the work on binder paper you can include that in your Rainbow Packet after this worksheet. If we end up with extra class time then portions of this may turn into required work. If that happens you will be told which problems are turned into required. Remember there is tons of other extra practice on the class website...and the entire internet! See me if you need help finding practice on a topic you are struggling with.

- Show work for ANY math problem.
- Include ALL units.
- Use a SINGLE DIMENSIONAL ANALYSIS line method set ups for ALL conversions.



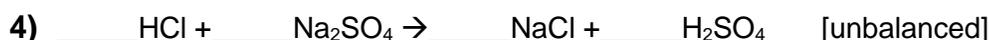
If you start with 10 g of lithium hydroxide, how many grams of lithium bromide will be produced?



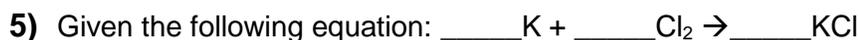
If you start with 45 grams of ethylene (C₂H₄), how many grams of carbon dioxide will be produced?



If you start with 5.5 grams of lithium chloride, how many grams of calcium chloride will be produced?



If you start with 20 grams of hydrochloric acid, how many grams of sulfuric acid will be produced?



How many grams of KCl is produced from 2.50 g of K and excess Cl₂. From 1.00 g of Cl₂?

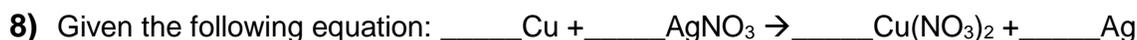


How many grams of NaOH is produced from 1.20 x 10² grams of Na₂O?

How many grams of Na₂O are required to produce 1.60 x 10² grams of NaOH?



What mass of iron is needed to react with 16.0 grams of sulfur? How many grams of FeS are produced?



How many moles of Cu are needed to react with 3.50 moles of AgNO₃?

If 89.5 grams of Ag were produced, how many grams of Cu reacted?

9) Molten iron and carbon monoxide are produced in a blast furnace by the reaction of iron(III) oxide and coke (pure carbon). The reaction is: _____ Fe₂O₃ + _____ C → _____ Fe + _____ CO

If 25.0 kilograms of pure Fe₂O₃ is used, how many kilograms of iron can be produced?

10) Laughing gas, N₂O can be turned into smog NO₂ by heating the laughing gas in the presence of oxygen. _____ N₂O + _____ O₂ → _____ NO₂

If 9.00 grams of laughing gas react. How many moles of oxygen react?

If 7.50 grams of oxygen react, how many grams of smog are produced?

Dougherty Valley HS Chemistry

Stoichiometry – Extra Stoichiometry Practice

- 11) The acid in your stomach that aids in breaking down proteins is called hydrochloric acid, HCl. Occasionally the glands that produce the HCl make more than is needed and you get those nasty sour burps. You can neutralize the excess acid by taking milk of magnesia. The active ingredient in milk of magnesia is $\text{Mg}(\text{OH})_2$.
The chemical reaction is. $\text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$
If you drink 2.00 grams of $\text{Mg}(\text{OH})_2$, how many moles of HCl are neutralized by reacting with the $\text{Mg}(\text{OH})_2$ to make your products?
- 12) Silver sulfide (Ag_2S) is the common tarnish on silver objects.
What weight of silver sulfide can be made from 1.23 mg of hydrogen sulfide (H_2S) obtained from a rotten egg? The reaction of formation of silver sulfide is:
 $\text{Ag}(\text{s}) + \text{H}_2\text{S}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{Ag}_2\text{S}(\text{s}) + \text{H}_2\text{O}(\text{l})$
- 13) A somewhat antiquated method for preparing chlorine gas involves heating hydrochloric acid with pyrolusite (manganese dioxide), a common manganese ore. (Reaction given below.)
How many kg of HCl react with 5.69 kg of manganese dioxide?
 $\text{HCl}(\text{aq}) + \text{MnO}_2(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{MnCl}_2(\text{aq}) + \text{Cl}_2(\text{g})$
- 14) Molten iron and carbon monoxide are produced in a blast furnace by the reaction of iron(III) oxide and coke (pure carbon). If 25.0 kilograms of pure Fe_2O_3 is used, how many kilograms of iron can be produced? The reaction is:
 $\text{Fe}_2\text{O}_3 + \text{C} \rightarrow \text{Fe} + \text{CO}$
- 15) 20.0 g of silver (I) nitrate is reacted with an excess of sodium chloride to produce silver(I) chloride. What mass of silver(I) chloride is produced?
 $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$ [Balanced]
- 16) 60.4 g of HCL is mixed with 56.1 g NaOH to produce water and table salt.
What mass of NaCl is produced?
 $\text{HCl} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCl}$ [Balanced]
- 17) How many grams of chlorine can be liberated from the decomposition of 64.0 g. of AuCl_3 by this reaction: $\text{AuCl}_3 \rightarrow \text{Au} + \text{Cl}_2$ [Unbalanced]
- 18) Calculate the mass of AgCl that can be prepared from 200. g of AlCl_3 and sufficient AgNO_3 , using this equation: $\text{AgNO}_3 + \text{AlCl}_3 \rightarrow \text{AgCl} + \text{Al}(\text{NO}_3)_3$ [Unbalanced]
- 19) Given this equation: $\text{KI} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbI}_2 + \text{KNO}_3$ [Unbalanced]
Calculate mass of PbI_2 produced by reacting of 30.0 g KI with excess $\text{Pb}(\text{NO}_3)_2$
- 20) How many grams of AuCl_3 can be made from 100.0 grams of chlorine by this reaction: $\text{Au} + \text{Cl}_2 \rightarrow \text{AuCl}_3$ [Unbalanced]
- 21) How many grams of Na are required to react completely with 75.0 grams of chlorine using this reaction: $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$ [Unbalanced]
- 22) Propane (C_3H_8) is burned. If I start with 5.0 grams of C_3H_8 , what is my theoretical yield of water? I got a percent yield of 75% How many grams of water did I make?
- 23) Beryllium + hydrochloric Acid. My theoretical yield of beryllium chloride was 10.7 grams. If my actual yield was 4.5 grams, what was my percent yield?

Dougherty Valley HS Chemistry

Stoichiometry – Extra Stoichiometry Practice

- 24) Sodium Chloride + Calcium Oxide. What is my theoretical yield of sodium oxide if I start with 20.0 grams of calcium oxide?
- 25) Iron (II) Bromide + Potassium Chloride. What is my theoretical yield of iron (II) chloride if I start with 34.0 grams of iron (II) bromide? What is my percent yield of iron (II) chloride if my actual yield is 4.0 grams?
- 26) $\text{TiS} + \text{H}_2\text{O} \rightarrow$
What is my percent yield of titanium (II) oxide if I start with 20.0 grams of titanium (II) sulfide and my actual yield of titanium (II) oxide is 22.0 grams?
- 27) $\text{U} + \text{Br}_2 \rightarrow \text{UBr}_6$ [unbalanced]
What is my actual yield of uranium hexabromide if I start with 100.0 grams of uranium and get a percent yield of 83%?
- 28) $\text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{SO}_3$ [Balanced]
If I start with 89 grams of sulfuric acid and produce 7.1 grams of water, what is my percent yield?